



STIC Search Report

EIC 1700

STIC Database Tracking Number: 132439

TO: Melvyn Andrews
Location: REM6A01
Art Unit : 1742
September 20, 2004

Case Serial Number: 10/031813

From: Kathleen Fuller
Location: EIC 1700
REMSSEN 4B28
Phone: 571/272-2505
Kathleen.Fuller@uspto.gov

Search Notes

=> FILE WPIX
FILE 'WPIX' ENTERED AT 11:55:43 ON 20 SEP 2004
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FILE LAST UPDATED: 15 SEP 2004 <20040915/UP>
MOST RECENT DERWENT UPDATE: 200459 <200459/DW>
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=> D QUE L49

L22 1 SEA FILE=REGISTRY ABB=ON 12634-54-7
L25 1 SEA FILE=REGISTRY ABB=ON MAGNESIUM/CN
L34 1543680 SEA FILE=HCAPLUS ABB=ON L22 OR L25 OR (MG OR MAGNESIUM)
L39 6125 SEA FILE=WPIX ABB=ON L34(6A) (MELT? OR MOLTEN? OR LIQ?)
L42 3 SEA FILE=WPIX ABB=ON L39 AND (DIFLUOROMETHANE OR PENTAFLUOROET
HANE OR DIFLUOROETHANE OR HEPTAFLUOROPROPANE OR (METHOXY OR
ETHOXY) (W)NONAFLUOROBUTANE OR DIHYDRODECAFLUOROPENTANE OR
TETRAFLUOROETHANE)
L43 0 SEA FILE=WPIX ABB=ON L39 AND (HYDROFLUOROETHER# OR HYDRO(W)FLU
ORO(W)ETHER# OR HYDROFLUORO(W)ETHER#)
L44 3 SEA FILE=WPIX ABB=ON L42 OR L43
L46 1 SEA FILE=COMPENDEX ABB=ON L39 AND HYDROFLUORO?
L49 4 SEA FILE=WPIX ABB=ON L44 OR L46

=> D L49 FULL 1-4

L49 ANSWER 1 OF 4 WPIX COPYRIGHT 2004 THOMSON DERWENT on STN
AN 2004-227174 [21] WPIX
DNN N2004-179528 DNC C2004-089559
TI Production of **magnesium** alloy strip comprises passing
molten alloy to feeding device, feeding molten alloy to chamber,
rotating rolls in which alloy is drawn from chamber, and flowing coolant
fluid through each roll.
DC M22 P53
IN ALLEN, R V; BORBRIDGE, W; EAST, D R; LIANG, D D
PA (CSIR) COMMONWEALTH SCI & IND RES ORG
CYC 105
PI WO 2004020126 A1 20040311 (200421)* EN 33 B22D011-06
RW: AT BE BG CH CY CZ DE DK EA EE ES FI FR GB GH GM GR HU IE IT KE LS
LU MC MW MZ NL OA PT RO SD SE SI SK SL SZ TR TZ UG ZM ZW

W: AE AG AL AM AT AU AZ BA BB BG BR BY BZ CA CH CN CO CR CU CZ DE DK
DM DZ EC EE ES FI GB GD GE GH GM HR HU ID IL IN IS JP KE KG KP KR
KZ LC LK LR LS LT LU LV MA MD MG MK MN MW MX MZ NI NO NZ OM PG PH
PL PT RO RU SC SD SE SG SK SL SY TJ TM TN TR TT TZ UA UG US UZ VC
VN YU ZA ZM ZW

ADT WO 2004020126 A1 WO 2003-AU1097 20030827

PRAI AU 2002-951075 20020829

IC ICM B22D011-06

AB WO2004020126 A UPAB: 20040326

NOVELTY - **Magnesium** alloy strip is produced by passing **molten** alloy from source of supply to feeding device, feeding molten alloy from feeding device through a nozzle (30) to a chamber (44), rotating rolls (32ab) in opposite direction in which alloy is drawn from the chamber, and flowing coolant fluid through each roll to provide internal cooling of the rolls.

DETAILED DESCRIPTION - Production of **magnesium** alloy strip involves passing **molten** alloy from source of supply to feeding device; feeding the molten alloy from the feeding device through the nozzle to the chamber formed between an elongate outlet of the nozzle and a pair of parallel rolls that are spaced one above the other to define a bite; rotating the rolls in opposite direction in which alloy is drawn from the chamber through the bite; and flowing coolant fluid through each roll to provide internal cooling of the rolls to cool the alloy received in the chamber by heat energy extraction by the cooled rolls to complete solidification of the alloy. The production method further includes maintaining the alloy held at the source at a temperature to maintain the alloy in the feed device at superheated temperature above its liquidus temperature, maintaining a depth of molten alloy in the feed device at a controlled constant height of 5-22 mm above a centerline of the bite in a plane containing the axes of the rolls, and maintaining heat energy extraction by the cooled rolls at a level to maintain alloy strip issuing from the bite at a surface temperature below 400 deg. C.

USE - Used in the production of magnesium alloy strip.

ADVANTAGE - Provides magnesium alloy strip of a required thickness and width.

DESCRIPTION OF DRAWING(S) - The figure shows an enlarged scale details relating to magnesium alloy solidification.

Nozzle 30

Rolls 32ab

Nip 34

Horizontal plates 36,37

Outlet 42

Chamber 44

Dwg.9/12

TECH WO 2004020126 A1UPTX: 20040326

TECHNOLOGY FOCUS - METALLURGY - Preferred Methods: The alloy at the source is held at 15-60 degrees C above its liquidus temperature. The level of heat extraction is able to maintain the surface temperature at 180-300 degrees C. The rolls apply a specific load of 2-500 kg/mm of roll length to the solidified alloy passing through the bite. A protective atmosphere is maintained over the molten alloy to safeguard against oxidation and risk of fire. The atmosphere includes a minor proportion of 1,1,1,2-tetrafluoroethane.

FS CPI GMPI

FA AB; GI

MC CPI: M22-G03A1A

L49 ANSWER 2 OF 4 WPIX COPYRIGHT 2004 THOMSON DERWENT on STN

AN 2002-055485 [07] WPIX

DNN N2002-040877 DNC C2002-015901
 TI Treatment of **molten magnesium** to protect
magnesium from reacting with air involves exposing **molten**
magnesium to gaseous mixture comprising fluorocarbon that includes
 limited amounts of hydrogen.
 DC E16 M25 P53
 IN MILBRATH, D S; OWENS, J G
 PA (MINN) 3M INNOVATIVE PROPERTIES CO
 CYC 96
 PI WO 2001083836 A2 20011108 (200207)* EN 26 C22B026-22
 RW: AT BE CH CY DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW MZ
 NL OA PT SD SE SL SZ TR TZ UG ZW
 W: AE AG AL AM AT AU AZ BA BB BG BR BY BZ CA CH CN CO CR CU CZ DE DK
 DM DZ EE ES FI GB GD GE GH GM HR HU ID IL IN IS JP KE KG KP KR KZ
 LC LK LR LS LT LU LV MA MD MG MK MN MW MX MZ NO NZ PL PT RO RU SD
 SE SG SI SK SL TJ TM TR TT TZ UA UG UZ VN YU ZA ZW
 AU 2001052990 A 20011112 (200222)
 NO 2002000027 A 20020207 (200227) C22B000-00
 KR 2002065466 A 20020813 (200309) C22B026-22
 EP 1278897 A2 20030129 (200310) EN C22B026-22
 R: AL AT BE CH CY DE DK ES FI FR GB GR IE IT LI LT LU LV MC MK NL PT
 RO SE SI TR
 US 2003034094 A1 20030220 (200316) C23C014-12
 US 6537346 B2 20030325 (200325) C22B009-05
 US 2003164068 A1 20030904 (200359) A62D001-02
 US 2003164069 A1 20030904 (200359) C22B009-00
 MX 2001013455 A1 20020801 (200367) C22B026-22
 JP 2003531965 W 20031028 (200373) 37 C22B026-22
 US 6685764 B2 20040203 (200413) C22B009-05
 US 6780220 B2 20040824 (200457) C22B026-22
 ADT WO 2001083836 A2 WO 2001-US9899 20010328; AU 2001052990 A AU 2001-52990
 20010328; NO 2002000027 A WO 2001-US9899 20010328, NO 2002-27 20020103; KR
 2002065466 A KR 2002-700054 20020103; EP 1278897 A2 EP 2001-926456
 20010328, WO 2001-US9899 20010328; US 2003034094 A1 Provisional US
 2000-202169P 20000504, US 2001-780256 20010209; US 6537346 B2 Provisional
 US 2000-202169P 20000504, US 2001-780256 20010209; US 2003164068 A1
 Provisional US 2000-202169P 20000504, CIP of US 2001-780256 20010209, US
 2003-394852 20030322; US 2003164069 A1 Provisional US 2000-202169P
 20000504, CIP of US 2001-780256 20010209, US 2003-394853 20030322; MX
 2001013455 A1 WO 2001-US9899 20010328, MX 2001-13455 20011219; JP
 2003531965 W JP 2001-580443 20010328, WO 2001-US9899 20010328; US 6685764
 B2-Provisional US 2000-202169P 20000504, CIP of US 2001-780256 20010209,
 US 2003-394852 20030322; US 6780220 B2 Provisional US 2000-202169P
 20000504, CIP of US 2001-780256 20010209, US 2003-394853 20030322
 FDT AU 2001052990 A Based on WO 2001083836; EP 1278897 A2 Based on WO
 2001083836; US 2003164068 A1 CIP of US 6537346; US 2003164069 A1 CIP of US
 6537346; MX 2001013455 A1 Based on WO 2001083836; JP 2003531965 W Based on
 WO 2001083836; US 6685764 B2 CIP of US 6537346; US 6780220 B2 CIP of US
 6537346
 PRAI US 2001-780256 20010209; US 2000-202169P 20000504;
 US 2003-394852 20030322; US 2003-394853 20030322
 IC ICM A62D001-02; C22B000-00; C22B009-00; C22B009-05; C22B026-22;
 C23C014-12
 ICS B22D001-00; B22D021-02; B22D021-04; C07C049-16; C22B009-02;
 C23C008-06; C23C026-00; G06F017-60
 AB WO 200183836 A UPAB: 20020130
 NOVELTY - **Molten magnesium** is treated by providing
molten magnesium and exposing the **magnesium** to
 a gaseous mixture comprising fluorocarbon such as perfluoroketone, and/or

hydrofluoroketone.

DETAILED DESCRIPTION - INDEPENDENT CLAIMS are also included for:

(A) solid magnesium including a film formed on the surface of the magnesium including a reaction product of molten magnesium and a fluorocarbon; and

(B) a method of extinguishing fire on the surface of magnesium.

USE - For protecting magnesium used in automotive and aerospace industries, from reacting with oxygen in air.

ADVANTAGE - The global warming potentials of PFKs and HFKs are quite low (less than 22200). Thus, the process is more environmentally friendly. PFKs and HFKs also react more fully with molten magnesium than does sulfur hexafluoride. As a result less unreacted cover gas can be emitted to the atmosphere and/or less cover gas can be required to produce a comparably performing protective film. As a result, useful concentrations of the cover gas can be lowered, reducing the global warming impact. The full substitution of the invention for SF6 can be accomplished without increasing the risk of worker safety since PFKs and HFs are of low toxicity, are non-flammable and are very innocuous materials. Surface films produced with the PFKs and HFKs are more stable to higher temperatures than those formed with sulfur dioxide, enabling work with higher melt temperatures (for additional alloys and more complex casting parts).

Dwg.0/0

TECH WO 200183836 A2UPTX: 20020130

TECHNOLOGY FOCUS - METALLURGY - Preferred Component: The gaseous mixture contains a carrier gas such as air, carbon dioxide, argon, and/or nitrogen.

Preferred Material: The magnesium is in the form of ingots or castings.

TECHNOLOGY FOCUS - ORGANIC CHEMISTRY - Preferred Compounds: The perfluoroketone (PFK) is 1,1,1,2,4,4,5,5,5-nonafluoro-2-trifluoromethyl-pentan-3-one, 1,1,1,2,4,5,5,5-octafluoro-2,4 di(trifluoromethyl)-pentan-3-one, 1,1,1,2,4,4,5,5,6,6,6-undecafluoro-2- trifluoromethyl-hexan-3-one, 1,1,1,2,4,4,5,5,6,6,7,7,7-tridecafluoro-2-trifluoromethyl-heptan-3-one, 1,1,1,3,3,4,4,5,5,6,6,7,7,8,8,8-octafluoro-octan-2-one, 1,1,1,2,2,4,4,5,5,6,6,6-dodecafluoro-hexan-3-one, 1,1,1,3,4,4,4-heptafluoro-3-trifluoromethyl-butan-2-one, perfluorocyclohexanone and their mixtures.

The hydrofluoroketone (HFK) is HCF2CF2C(O)CF(CF3)2, CF3C(O)CH2C(O)CF3, C2H5C(O)CF(CF3)2, CF2CF2C(O)CH3, (CF3)2CFC(O)CH3, CF3CF2C(O)CHF2, CF3CF2C(O)CH2F, CF3CF2C(O)CH2CF3, CF3CF2C(O)CH2CH3, CF3CF2C(O)CH2CHF2, CF3CF2C(O)CH2CHF2, CF3CF2C(O)CH2CH2F, CF3CF2C(O)CHFCH3, CF3CF2C(O)CHFCHF2, CF3CF2C(O)CF2CH2F, (CF3)2CFC(O)CHF2, (CF3)2CFC(O)CH2F, CF3CF(CH2F)C(O)CHF2, CH3CF(CH2F)C(O)CH2F, and/or CF3CF(CH2F)C(O)CF3, and their mixtures..

ABEX WO 200183836 A2UPTX: 20020130

EXAMPLE - Pure magnesium (3 kg) was heated to 680 degreesC and two cover gases were continuously applied separately to the surface of the molten magnesium. The invented method used 1,1,1,2,4,4,5,5,5-nonafluoro-2-trifluoromethyl-pentan-3-one (perfluorinated ketone) as cover gas, while the comparative process used methoxy nonafluorobutane. The cover gases were evaluated at 1% and at decreasing volumetric concentrations in air. At all concentrations, the perfluorinated ketone produced a film visually thinner and more elastic than that produced by the comparative process. The film produced was stable and did not change appearance over at least 30 minutes. This was in contrast to the test using methoxy nonafluorobutane, where metal burning was noted when the cover gas concentration was reduced to 625 ppm.

FS CPI GMPI

FA AB; DCN
MC CPI: E10-F02A3; E10-F02C; M25-G16

L49 ANSWER 3 OF 4 WPIX COPYRIGHT 2004 THOMSON DERWENT on STN
AN 2001-070604 [08] WPIX
DNN N2001-053480 DNC C2001-019551
TI Cover gas composition for protecting **molten magnesium**
(alloy) comprises fluorine containing inhibiting agent and carrier gas.
DC E16 M22 P35 P53
IN BAKER, P W; CASHION, S P; FROST, M T; KORN, C J; RICKETTS, N J
PA (CAST-N) CAST CENT PTY LTD
CYC 94

applicants

PI WO 2000064614 A1 20001102 (200108)* EN 20 B22D021-04
RW: AT BE CH CY DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW NL
OA PT SD SE SL SZ TZ UG ZW
W: AE AG AL AM AT AU AZ BA BB BG BR BY CA CH CN CR CU CZ DE DK DM DZ
EE ES FI GB GD GE GH GM HR HU ID IL IN IS JP KE KG KP KR KZ LC LK
LR LS LT LU LV MA MD MG MK MN MW MX NO NZ PL PT RO RU SD SE SG SI
SK SL TJ TM TR TT TZ UA UG US UZ VN YU ZA ZW

AU 2000040930 A 20001110 (200109)
BR 2000010137 A 20020122 (200216) B22D021-04
NO 2001005264 A 20011221 (200221) B22D000-00
EP 1204499 A1 20020515 (200239) EN B22D021-04
R: AL AT BE CH CY DE DK ES FI FR GB GR IE IT LI LT LU LV MC MK NL PT
RO SE SI

CZ 2001003817 A3 20020515 (200241) B22D021-04
KR 2002011397 A 20020208 (200255) B22D021-04
HU 2002000990 A2 20020729 (200258) B22D021-04
CN 1352583 A 20020605 (200261) B22D021-04
SK 2001001556 A3 20020806 (200261) B22D021-04
JP 2002541999 W 20021210 (200301) 18 A62C003-06
ZA 2001008862 A 20021224 (200309) 24 B22D000-00
TW 500805 A 20020901 (200334) C22B026-22
NZ 515084 A 20031031 (200380) B22D021-04
AU 766844 B 20031023 (200381) B22D021-04
MX 2001010941 A1 20030601 (200417) A62D001-02

ADT WO 2000064614 A1 WO 2000-AU393 20000428; AU 2000040930 A AU 2000-40930
20000428; BR 2000010137 A BR 2000-10137 20000428; WO 2000-AU393 20000428;
NO 2001005264 A WO 2000-AU393 20000428; NO 2001-5264 20011026; EP 1204499
A1 EP 2000-920274 20000428; WO 2000-AU393 20000428; CZ 2001003817 A3 WO
2000-AU393 20000428; CZ 2001-3817 20000428; KR 2002011397 A KR 2001-713855
20011029; HU 2002000990 A2 WO 2000-AU393 20000428; HU 2002-990 20000428;
CN 1352583 A CN 2000-808146 20000428; SK 2001001556 A3 WO 2000-AU393
20000428; SK 2001-1556 20000428; JP 2002541999 W JP 2000-613596 20000428;
WO 2000-AU393 20000428; ZA 2001008862 A ZA 2001-8862 20011026; TW 500805 A
TW 2000-108130 20000504; NZ 515084 A NZ 2000-515084 20000428; WO
2000-AU393 20000428; AU 766844 B AU 2000-40930 20000428; MX 2001010941 A1
WO 2000-AU393 20000428; MX 2001-10941 20011026

FDT AU 2000040930 A Based on WO 2000064614; BR 2000010137 A Based on WO
2000064614; EP 1204499 A1 Based on WO 2000064614; CZ 2001003817 A3 Based
on WO 2000064614; HU 2002000990 A2 Based on WO 2000064614; SK 2001001556
A3 Based on WO 2000064614; JP 2002541999 W Based on WO 2000064614; NZ
515084 A Based on WO 2000064614; AU 766844 B Previous Publ. AU 2000040930,
Based on WO 2000064614; MX 2001010941 A1 Based on WO 2000064614

PRAI AU 1999-15 19990428
IC ICM A62C003-06; A62D001-02; B22D000-00
ICS B22D021-02; C22B004-02; C22C001-02; C22C023-00
ICA B22D021-04; C22B026-22
AB WO 200064614 A UPAB: 20010207

NOVELTY - Cover gas composition comprises a fluorine containing inhibiting agent and a carrier gas. Each component has a Global Warming Potential (GWP) of less than 5000.

DETAILED DESCRIPTION - INDEPENDENT CLAIMS are also included for:

- (a) a method of protecting **molten magnesium** (alloy);
- (b) use of the inhibiting agent for preventing or minimizing oxidation of **molten magnesium** (alloy);
- (c) a method of extinguishing a magnesium (alloy) fire.

USE - For use as cover gases for protecting **molten magnesium** (alloys).

ADVANTAGE - The inventive cover gas composition provides good industrial scale protection of **molten magnesium** at a lower concentration.

Dwg.0/0

TECH WO 200064614 A1UPTX: 20010207

TECHNOLOGY FOCUS - INORGANIC CHEMISTRY - Preferred Gas: The carrier gas is carbon dioxide, argon, nitrogen, and/or preferably dry air.

TECHNOLOGY FOCUS - ORGANIC CHEMISTRY - Preferred Materials: The inhibiting agent is **difluoromethane, pentafluoromethane, difluoroethane, heptafluoropropane, methoxy-nonafluorobutane, ethoxy-nonafluorobutane, dihydrodecafluoropentane**, and/or preferably 1,1,1,2-tetrafluoroethane.

Preferred Properties: The inhibiting agent has a boiling point of less than 100 degrees C and has no ozone depletion potential.

The composition has a GWP of less than 3000 (preferably less than 1500).

Preferred Composition: The composition contains inhibiting agent (less than 1, preferably less than 0.1 vol.%).

ABEX WO 200064614 A1UPTX: 20010207

EXAMPLE - A crucible furnace containing 100g of molten pure magnesium at 680 degrees C was blanketed with a gaseous composition consisting of 1,1,1,2-tetrafluoroethane (0.02 volume%) and dry air (balance). Good molten protection was observed, with the formation of a thin protective surface film.

FS CPI GMPI

FA AB; DCN

MC CPI: E10-H04A3; E31-H03; E31-J; E31-N05C; M22-G

L49 ANSWER 4 OF 4 WPIX COPYRIGHT 2004 THOMSON DERWENT on STN

AN 1996-171062 [17] WPIX

DNC C1996-053941

TI Catalyst preparation for production of 1,1,1,2-tetra fluoroethane - by producing

admixture of compsn. comprising chromium hydroxide hydrate and e.g. magnesium chloride with aqueous metal salt, reacting in liquid phase, etc..

DC E16 J04

IN KIM, H; LEE, B; PARK, K; KIM, H S; LEE, B G; PARK, K Y

PA (KOAD) KOREA INST SCI & TECHNOLOGY; (KOAD) KOREA ADV INST SCI & TECHNOLOGY
CYC 2

PI US 5500400 A 19960319 (199617)* 4 B01J023-26

KR 9616683 B1 19961220 (199931) B01J023-26

ADT US 5500400 A US 1994-302636 19940908; KR 9616683 B1 KR 1993-27091 19931209

PRAI KR 1993-27091 19931209

IC ICM B01J023-26

ICS B01J023-78; C07C019-08

AB US 5500400 A UPAB: 19960428

The catalyst for the production of 1,1,1,2-tetrafluoroethane

(HFC-134a) is prepared by: (a) preparing an admixture of a compsn.

- comprising Cr hydroxide hydrate and MgCl₂ or CaCl₂ in a weight ratio of Cr to Mg or Ca of 1:0.3-10 with an aqueous metal salt solution from Ce, Mg, Ni and Al chlorides, (b) reacting the mixture in the liquid phase with an aqueous HF solution

to give a paste; and (c) sintering the paste.

Also claimed is a catalyst for the production of HFC-134a represented by formula Cr_aL_bM_cO_xF_y (I), where: L = Mg or Ca; M = one of Ce, Ni, Zn, and Al; and a, b, c, x and y = 0-2.

USE - Used in the production of HFC-134a for use in refrigerators, car cooling systems, etc., by reaction of HCFC-133a with HF.

ADVANTAGE - The catalyst has improved selectivity and durability, and remains active over extended periods. Pre-treatment with HF solution is not required, and the supply of O₂ to reduce catalyst deactivation is also not required, thus reducing by-prod. formation, HCl oxidation, and water formation, lowering corrosion in the appts..

Dwg.0/0

FS CPI

FA AB; DCN

MC CPI: E10-H03A3; E35-C; E35-P; E35-W; J04-E04; N01-B; N01-C; N03-A; N03-D01; N03-F; N06-E

=> FILE HCAPLU

FILE 'HCAPLUS' ENTERED AT 11:56:28 ON 20 SEP 2004

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FILE COVERS 1907 - 20 Sep 2004 VOL 141 ISS 13

FILE LAST UPDATED: 19 Sep 2004 (20040919/ED)

This file contains CAS Registry Numbers for easy and accurate substance identification.

=> D QUE L37

L22 1 SEA FILE=REGISTRY ABB=ON 12634-54-7
 L23 13 SEA FILE=REGISTRY ABB=ON (124-38-9/BI OR 12634-54-7/BI OR 138495-42-8/BI OR 163702-05-4/BI OR 219484-64-7/BI OR 33660-75-2/BI OR 354-33-6/BI OR 7439-95-4/BI OR 7440-37-1/BI OR 75-10-5/BI OR 75-37-6/BI OR 7727-37-9/BI OR 811-97-2/BI)
 L24 8 SEA FILE=REGISTRY ABB=ON L23 AND 1-20/F
 L25 1 SEA FILE=REGISTRY ABB=ON MAGNESIUM/CN
 L28 1 SEA FILE=REGISTRY ABB=ON "1,1,1,2-TETRAFLUOROETHANE"/CN
 L32 1 SEA FILE=REGISTRY ABB=ON 163702-08-7
 L34 1543680 SEA FILE=HCAPLUS ABB=ON L22 OR L25 OR (MG OR MAGNESIUM)
 L35 19560 SEA FILE=HCAPLUS ABB=ON L34(6A)(MELT? OR MOLTEN? OR LIQ?)
 L36 8793 SEA FILE=HCAPLUS ABB=ON L24 OR L28 OR L32

L37 8 SEA FILE=HCAPLUS ABB=ON L35 AND L36

=> FILE COMPENDEX

FILE 'COMPENDEX' ENTERED AT 11:56:41 ON 20 SEP 2004

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FILE COVERS 1970 TO DATE.

<<< SIMULTANEOUS LEFT AND RIGHT TRUNCATION AVAILABLE IN
THE BASIC INDEX >>>

=> D QUE L47

L22 1 SEA FILE=REGISTRY ABB=ON 12634-54-7
L25 1 SEA FILE=REGISTRY ABB=ON MAGNESIUM/CN
L34 1543680 SEA FILE=HCAPLUS ABB=ON L22 OR L25 OR (MG OR MAGNESIUM)
L39 6125 SEA FILE=WPIX ABB=ON L34 (6A) (MELT? OR MOLTEN? OR LIQ?)
L42 3 SEA FILE=WPIX ABB=ON L39 AND (DIFLUOROMETHANE OR PENTAFLUOROET
HANE OR DIFLUOROETHANE OR HEPTAFLUOROPROPANE OR (METHOXY OR
ETHOXY) (W)NONAFLUOROBUTANE OR DIHYDRODECAFLUOROPENTANE OR
TETRAFLUROETHANE)
L43 0 SEA FILE=WPIX ABB=ON L39 AND (HYDROFLUROETHER# OR HYDRO(W) FLU
ORO(W)ETHER# OR HYDROFLURO(W)ETHER#)
L45 1 SEA FILE=COMPENDEX ABB=ON L42 OR L43
L46 1 SEA FILE=COMPENDEX ABB=ON L39 AND HYDROFLURO?
L47 1 SEA FILE=COMPENDEX ABB=ON L45 OR L46

=> FILE METADEX

FILE 'METADEX' ENTERED AT 11:57:02 ON 20 SEP 2004

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FILE LAST UPDATED: 30 AUG 2004 <20040830/UP>

FILE COVERS 1966 TO DATE.

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=> D QUE L50

L22 1 SEA FILE=REGISTRY ABB=ON 12634-54-7
L25 1 SEA FILE=REGISTRY ABB=ON MAGNESIUM/CN
L34 1543680 SEA FILE=HCAPLUS ABB=ON L22 OR L25 OR (MG OR MAGNESIUM)
L39 6125 SEA FILE=WPIX ABB=ON L34 (6A) (MELT? OR MOLTEN? OR LIQ?)
L42 3 SEA FILE=WPIX ABB=ON L39 AND (DIFLUOROMETHANE OR PENTAFLUROET
HANE OR DIFLUOROETHANE OR HEPTAFLUROPROPANE OR (METHOXY OR
ETHOXY) (W)NONAFLUROOBUTANE OR DIHYDRODECAFLUROOPENTANE OR
TETRAFLUROETHANE)
L43 0 SEA FILE=WPIX ABB=ON L39 AND (HYDROFLUROETHER# OR HYDRO(W) FLU
ORO(W)ETHER# OR HYDROFLURO(W)ETHER#)
L44 3 SEA FILE=WPIX ABB=ON L42 OR L43
L46 1 SEA FILE=COMPENDEX ABB=ON L39 AND HYDROFLURO?
L50 3 SEA FILE=METADEX ABB=ON L44 OR L46

=> FILE NTIS

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FILE LAST UPDATED: 18 SEP 2004 <20040918/UP>
FILE COVERS 1964 TO DATE.

<<<SIMULTANEOUS LEFT AND RIGHT TRUNCATION AVAILABLE IN
THE BASIC INDEX (/BI) >>>

=> D QUE L51

L22 1 SEA FILE=REGISTRY ABB=ON 12634-54-7
L25 1 SEA FILE=REGISTRY ABB=ON MAGNESIUM/CN
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L39 6125 SEA FILE=WPIX ABB=ON L34 (6A) (MELT? OR MOLTEN? OR LIQ?)
L42 3 SEA FILE=WPIX ABB=ON L39 AND (DIFLUOROMETHANE OR PENTAFLUOROET
HANE OR DIFLUOROETHANE OR HEPTAFLUOROPROPANE OR (METHOXY OR
ETHOXY) (W)NONAFLUOROBUTANE OR DIHYDRODECAFLUOROPENTANE OR
TETRAFLUOROETHANE)
L43 0 SEA FILE=WPIX ABB=ON L39 AND (HYDROFLUOROETHER# OR HYDRO(W) FLU
ORO(W)ETHER# OR HYDROFLUORO(W)ETHER#)
L44 3 SEA FILE=WPIX ABB=ON L42 OR L43
L46 1 SEA FILE=COMPENDEX ABB=ON L39 AND HYDROFLUORO?
L51 0 SEA FILE=NTIS ABB=ON L44 OR L46

=> FILE JICST

FILE 'JICST-EPLUS' ENTERED AT 11:57:42 ON 20 SEP 2004
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FILE COVERS 1985 TO 13 SEP 2004 (20040913/ED)

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=> D QUE L52

L22 1 SEA FILE=REGISTRY ABB=ON 12634-54-7
L25 1 SEA FILE=REGISTRY ABB=ON MAGNESIUM/CN
L34 1543680 SEA FILE=HCAPLUS ABB=ON L22 OR L25 OR (MG OR MAGNESIUM)
L39 6125 SEA FILE=WPIX ABB=ON L34 (6A) (MELT? OR MOLTEN? OR LIQ?)
L42 3 SEA FILE=WPIX ABB=ON L39 AND (DIFLUOROMETHANE OR PENTAFLUOROET
HANE OR DIFLUOROETHANE OR HEPTAFLUOROPROPANE OR (METHOXY OR
ETHOXY) (W)NONAFLUOROBUTANE OR DIHYDRODECAFLUOROPENTANE OR
TETRAFLUOROETHANE)
L43 0 SEA FILE=WPIX ABB=ON L39 AND (HYDROFLUOROETHER# OR HYDRO(W) FLU
ORO(W)ETHER# OR HYDROFLUORO(W)ETHER#)
L44 3 SEA FILE=WPIX ABB=ON L42 OR L43
L46 1 SEA FILE=COMPENDEX ABB=ON L39 AND HYDROFLUORO?
L52 0 SEA FILE=JICST-EPLUS ABB=ON L44 OR L46

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FILE 'JAPIO' ENTERED AT 11:57:55 ON 20 SEP 2004
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FILE LAST UPDATED: 3 SEP 2004 <20040903/UP>
FILE COVERS APR 1973 TO APRIL 30, 2004

<<< GRAPHIC IMAGES AVAILABLE >>>

KATHLEEN FULLER EIC 1700 REMSEN 4B28 571/272-2505

=> D QUE L53

L22 1 SEA FILE=REGISTRY ABB=ON 12634-54-7
 L25 1 SEA FILE=REGISTRY ABB=ON MAGNESIUM/CN
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 L39 6125 SEA FILE=WPIX ABB=ON L34(6A)(MELT? OR MOLTEN? OR LIQ?)
 L42 3 SEA FILE=WPIX ABB=ON L39 AND (DIFLUOROMETHANE OR PENTAFLUOROETHANE OR DIFLUOROETHANE OR HEPTAFLUOROPROPANE OR (METHOXY OR ETHOXY)(W)NONAFLUOROBUTANE OR DIHYDRODECAFLUOROPENTANE OR TETRAFLUOROETHANE)
 L43 0 SEA FILE=WPIX ABB=ON L39 AND (HYDROFLUOROETHER# OR HYDRO(W)FLUORO(W)ETHER# OR HYDROFLUORO(W)ETHER#)
 L44 3 SEA FILE=WPIX ABB=ON L42 OR L43
 L46 1 SEA FILE=COMPENDEX ABB=ON L39 AND HYDROFLUORO?
 L53 0 SEA FILE=JAPIO ABB=ON L44 OR L46

=> DUP REM L37 L47 L50

FILE 'HCAPLUS' ENTERED AT 11:58:18 ON 20 SEP 2004
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FILE 'METADEX' ENTERED AT 11:58:18 ON 20 SEP 2004
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 PROCESSING COMPLETED FOR L37
 PROCESSING COMPLETED FOR L47
 PROCESSING COMPLETED FOR L50
 L54 9 DUP REM L37 L47 L50 (3 DUPLICATES REMOVED)

=> D L54 ALL HITSTR 1-9

L54 ANSWER 1 OF 9 HCAPLUS COPYRIGHT 2004 ACS on STN
 AN 2003:409780 HCAPLUS
 DN 139:10048
 ED Entered STN: 29 May 2003
 TI **Magnesium melt** protection at **Magnesium**
 Elektron using HFC-134a
 AU Lyon, Paul; Rogers, Philip D.; King, John F.; Cashion, Simon P.; Ricketts, Nigel J.
 CS Development Dept., Magnesium Elektron, Manchester, M27 8DD, UK
 SO Magnesium Technology 2003, Proceedings of the Symposium held during the 2003 TMS Annual Meeting, San Diego, CA, United States, ar. 2-6, 2003 (2003), 11-14. Editor(s): Kaplan, Howard I. Publisher: Minerals, Metals & Materials Society, Warrendale, Pa.
 CODEN: 69DYJV; ISBN: 0-87339-533-6
 DT Conference
 LA English
 CC 56-2 (Nonferrous Metals and Alloys)
 AB Several workers are actively seeking alternatives to SF6. CAST/AMC have patented the use of a HFC gas HFC-134a. This gas has a GWP 95% lower than SF6. Magnesium Elektron (MEL) and CAST have collaborated on the use of HFC-134a to achieve successful production plant trials for ingot manufacture
 This

paper provides details and results of those plant trials at Magnesium Elektron. Magnesium Elektron ingot production is currently based on either SF6 or SO2 as the active gas. Comparative evaluation of HFC-134a included, development of gas mixing equipment, distribution optimization, optimization of concentration and flow rates and assessment of breakdown products

and potential implications. The results, from over 150 production-scale **Mg-Al-Zn melts**, demonstrated that HFC-134a could offer equal protection to SF6 or SO2. Use of HFC-134a by Magnesium Elektron for com. production is proposed.

ST **magnesium** casting **melt** protection fluorinated hydrocarbon

IT **Melting**

(alloy; **magnesium melt** protection in casting with HFC-134a)

IT Hydrocarbons, uses

RL: NUU (Other use, unclassified); USES (Uses)

(fluoro; **magnesium melt** protection in casting with HFC-134a)

IT Casting of metals

Controlled atmospheres

(**magnesium melt** protection in casting with HFC-134a)

IT **811-97-2**, HFC-134a

RL: NUU (Other use, unclassified); USES (Uses)

(**magnesium melt** protection in casting with HFC-134a)

IT 135860-09-2, AZ91D

RL: PEP (Physical, engineering or chemical process); PYP (Physical process); PROC (Process)

(**magnesium melt** protection in casting with HFC-134a)

RE.CNT 9 THERE ARE 9 CITED REFERENCES AVAILABLE FOR THIS RECORD

RE

(1) Bartos, S; Magnesium Technology 2000 2000, P83

(2) Cashion, S; Paper presented at the 8th Annual IMA Magnesium in Automotive Seminar 2000

(3) Cashion, S; PhD Thesis, University of Queensland 1999

(4) Erickson, S; Foundry Management and Technology 1998, V126(6), P38

(5) Fruehling, J; PhD thesis, University of Michigan 1970

(6) Gjestland, H; The 3rd International Magnesium Conference 1996, P33

(7) Granier, C; Scientific Assessment of Ozone Depletion 1998

(8) Reimers, H; US 1972317 1934 HCAPLUS

(9) Ricketts, N; Magnesium Technology 2001 2001, P31 HCAPLUS

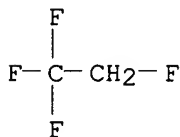
IT **811-97-2**, HFC-134a

RL: NUU (Other use, unclassified); USES (Uses)

(**magnesium melt** protection in casting with HFC-134a)

RN **811-97-2** HCAPLUS

CN Ethane, 1,1,1,2-tetrafluoro- (6CI, 7CI, 8CI, 9CI) (CA INDEX NAME)



L54 ANSWER 2 OF 9 HCAPLUS COPYRIGHT 2004 ACS on STN
 AN 2002:403070 HCAPLUS
 DN 137:220130
 ED Entered STN: 30 May 2002
 TI Characterisation of the surface films formed on **molten magnesium** in different protective atmospheres
 AU Pettersen, Gunnar; Ovrelid, E.; Tranell, G.; Fenstad, J.; Gjestland, H.
 CS SINTEF Materials Technology, Trondheim, N-7465, Norway
 SO Materials Science & Engineering, A: Structural Materials: Properties, Microstructure and Processing (2002), A332(1-2), 285-294
 CODEN: MSAPE3; ISSN: 0921-5093
 PB Elsevier Science B.V.
 DT Journal
 LA English
 CC 56-2 (Nonferrous Metals and Alloys)
 AB **Molten Mg** oxidizes rapidly during casting and handling unless it is protected by an atmospheric that stabilizes the surface. In this article, results from the anal. of **Mg melt** surfaces exposed to SO₂ and different fluorine-containing atmospheres are reported. The microstructure of the surface films, formed during controlled exposure in laboratory scale expts., have been characterized using X-ray diffraction (XRD), electron probe microanal. (EPMA) and transmission electron microscopy (TEM). Both SO₂ and the fluorine-containing gases were found to protect the melt from burning and vaporization in oxidizing atmospheres. The protected surfaces generally had a shiny metallic appearance, but turned dull gray after extended exposure to high concns. of fluorine containing gases. All the surface films initially consisted of small crystallites of MgO forming a thin continuous film. This film was found to contain some sulfur when the melt was protected by SO₂, while fluorine was the only element detected in the oxide when SF₆ or other fluorine-containing gases were used for protection. With increasing exposure time, the films gradually grew thicker and the fluorine/oxygen-ratio of the films formed in fluorine-containing atmospheres increased. Finally, after long term exposure to fluorine containing atmospheres, the thermodynamically stable MgF₂-phase was formed. In a N₂ atmosphere, SO₂ and SF₆-addns. did not protect the magnesium, indicating that a rapid initial formation of MgO is necessary to obtain protective films.
 ST controlled atm film formation **molten magnesium melt** oxidn prevention
 IT Casting of metals
 . Controlled atmospheres
 (characterization of the surface films formed on **molten magnesium** in different protective atmospheres)
 IT 7439-95-4P, Magnesium, preparation
 RL: IMF (Industrial manufacture); PRP (Properties); RCT (Reactant); TEM (Technical or engineered material use); PREP (Preparation); RACT (Reactant or reagent); USES (Uses)
 (characterization of the surface films formed on **molten magnesium** in different protective atmospheres)
 IT 7446-09-5, Sulfur dioxide, uses
 RL: TEM (Technical or engineered material use); USES (Uses)
 (characterization of the surface films formed on **molten magnesium** in different protective atmospheres)
 IT 75-28-5, Isobutane 76-19-7 **811-97-2** 454695-19-3, ICEON 49
 RL: NUU (Other use, unclassified); USES (Uses)
 (gas atmosphere containing; characterization of the surface films formed on **molten magnesium** in different protective atmospheres)
 IT 2551-62-4, Sulfur hexafluoride 7783-54-2, Trifluoroamine

RL: NUU (Other use, unclassified); USES (Uses)
 (gas atmosphere; characterization of the surface films formed on
molten magnesium in different protective atmospheres)
 IT 1309-48-4, Magnesium oxide (MgO), properties 7783-40-6, Magnesium
 fluoride (MgF₂) 12032-36-9, Magnesium sulfide (MgS)
 RL: FMU (Formation, unclassified); PRP (Properties); TEM (Technical or
 engineered material use); FORM (Formation, nonpreparative); USES (Uses)
 (protective phase; characterization of the surface films formed on
molten magnesium in different protective atmospheres)

RE.CNT 11 THERE ARE 11 CITED REFERENCES AVAILABLE FOR THIS RECORD
 RE

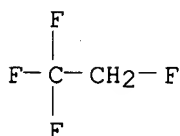
- (1) Cashion, S; PhD thesis, Department of Mining, Minerals and Materials
 Engineering, University of Queensland 1998
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- (3) Couling, S; Proc of the 36th Annual World Conference on Magnesium 1979, P54
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- (5) Gjestland, H; Proc of the Third Int Magnesium Conference 1996, P33
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 Sets 1-44
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 (ASST2000) 2000, P553 HCAPLUS
- (9) Putz, H; Computational Materials Science 1998, V11, P309 HCAPLUS
- (10) Reimers, H; US 1972317 1934 HCAPLUS
- (11) Roine, A; HSC Chemical Reaction and Equilibrium Software with extensive
 Thermochemical Database, version 2.03 1994

IT 811-97-2

RL: NUU (Other use, unclassified); USES (Uses)
 (gas atmosphere containing; characterization of the surface films formed on
molten magnesium in different protective atmospheres)

RN 811-97-2 HCAPLUS

CN Ethane, 1,1,1,2-tetrafluoro- (6CI, 7CI, 8CI, 9CI) (CA INDEX NAME)



L54 ANSWER 3 OF 9 HCAPLUS COPYRIGHT 2004 ACS on STN DUPLICATE 1

AN 2002:438031 HCAPLUS

DN 137:173739

ED Entered STN: 11 Jun 2002

TI Replacing SF₆ with the hydrofluorocarbon gas HFC-134a for
magnesium melt protection

AU Cashion, Simon P.; Ricketts, Nigel J.

CS Cooperative Research Centre for Cast Metals Manufacturing (CAST)
 Department of Mining, Minerals and Materials Engineering, The University
 of Queensland, St. Lucia, 4072, Australia

SO Greenhouse Gases in the Metallurgical Industries: Policies, Abatement and
 Treatment, Proceedings of the International Symposium on Greenhouse Gases
 in the Metallurgical Industries: Policies, Abatement and Treatment,
 Toronto, ON, Canada, Aug. 26-29, 2001 (2001), 315-324. Editor(s):
 Pickles, Chris A. Publisher: Canadian Institute of Mining, Metallurgy and
 Petroleum, Montreal, Que.

- CODEN: 69CROA; ISBN: 1-894475-15-1
 DT Conference
 LA English
 CC 59-2 (Air Pollution and Industrial Hygiene)
 Section cross-reference(s): 56
 AB Since the early 1970s, sulfur hexafluoride (SF6) has been used to protect **molten magnesium** from oxidation. However the use of SF6 is coming under increased scrutiny. It is now recognized as a very potent greenhouse gas, and has been targeted for phase-out under the Kyoto protocol. In the search for an alternative to SF6, the Cooperative Research Center for Cast Metals Manufacturing (CAST) in Australia is developing a cover gas system based upon the hydrofluorocarbon gas 1,1,1,2-tetrafluoroethane (HFC-134a). This gas has a global warming potential 17 times lower than SF6 and has been shown to be an effective replacement for SF6 in many applications. Recent experience with CAST's new cover gas system in laboratory and plant trials is discussed as well as the global implications of replacing SF6 with HFC-134a.
 ST air pollution greenhouse gas sulfur hexafluoride tetrafluoroethane magnesium metallurgy
 IT Air pollution
 Greenhouse gases
 (replacing the greenhouse gas SF6 with tetrafluoroethane for **magnesium melt** protection)
 IT 135860-09-2P, Az91d
 RL: IMF (Industrial manufacture); PREP (Preparation)
 (AZ91D, protection of molten; replacing the greenhouse gas SF6 with tetrafluoroethane for **magnesium melt** protection)
 IT 7664-39-3P, Hydrogen fluoride, biological studies
 RL: ADV (Adverse effect, including toxicity); BYP (Byproduct); BIOL (Biological study); PREP (Preparation)
 (decomposition product of HFC-134a; replacing the greenhouse gas SF6 with tetrafluoroethane for **magnesium melt** protection)
 IT 7439-95-4P, **Magnesium**, preparation
 RL: IMF (Industrial manufacture); PREP (Preparation)
 (protection of **molten**; replacing the greenhouse gas SF6 with tetrafluoroethane for **magnesium melt** protection)
 IT 811-97-2, Ethane, 1,1,1,2-tetrafluoro- 2551-62-4, Sulfur fluoride (SF6), (OC-6-11)-
 RL: NUU (Other use, unclassified); POL (Pollutant); OCCU (Occurrence); USES (Uses)
 (replacing the greenhouse gas SF6 with tetrafluoroethane for **magnesium melt** protection)

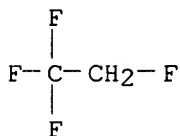
RE.CNT 13 THERE ARE 13 CITED REFERENCES AVAILABLE FOR THIS RECORD
 RE

- (1) Australian Bureau Of Statistics; Motor Vehicle Census 1998, 9309.0, P3
- (2) Bartos, S; Magnesium Technology 2000 2000, P83
- (3) Cashion, S; Paper presented at the 8th Annual IMA Magnesium in Automotive Seminar 2000
- (4) Cashion, S; Ph D Thesis, University of Queensland 1999
- (5) Erickson, S; Foundry Management and Technology 1998, V126(6), P38
- (6) Fruehling, J; PhD thesis, University of Michigan 1970
- (7) Gjestland, H; The 3rd International Magnesium Conference 1996, P33
- (8) Granier, C; Scientific Assessment of Ozone Depletion:1998
- (9) Mellerud, T; The 56th IMA Annual Meeting 1999
- (10) Reimers, H; US 1972317 1934 HCAPLUS
- (11) Ricketts, N; WO PCTAU0000393 2000
- (12) Ricketts, N; Magnesium Technology 2001 2001, P31 HCAPLUS
- (13) Wilkenfeld, G; Analysis of Trends and Greenhouse Indicators 1990 to 1998 2000, P41

IT 7439-95-4P, Magnesium, preparation
 RL: IMF (Industrial manufacture); PREP (Preparation)
 (protection of **molten**; replacing the greenhouse gas SF6 with
 tetrafluoroethane for **magnesium melt** protection)
 RN 7439-95-4 HCAPLUS
 CN Magnesium (8CI, 9CI) (CA INDEX NAME)

Mg

IT 811-97-2, Ethane, 1,1,1,2-tetrafluoro-
 RL: NUU (Other use, unclassified); POL (Pollutant); OCCU (Occurrence);
 USES (Uses)
 (replacing the greenhouse gas SF6 with tetrafluoroethane for
magnesium melt protection)
 RN 811-97-2 HCAPLUS
 CN Ethane, 1,1,1,2-tetrafluoro- (6CI, 7CI, 8CI, 9CI) (CA INDEX NAME)



L54 ANSWER 4 OF 9 HCAPLUS COPYRIGHT 2004 ACS on STN DUPLICATE 2
 AN 2001:272575 HCAPLUS
 DN 134:343042
 ED Entered STN: 18 Apr 2001
 TI Hydrofluorocarbons as a replacement for sulphur hexafluoride in magnesium
 processing
 AU Ricketts, Nigel J.; Cashion, Simon P.
 CS Cooperative Research Centre for Cast Metals Manufacturing (CAST),
 Pullenvale, 4069, Australia
 SO Magnesium Technology 2001, Proceedings of the Symposium held during the
 TMS Annual Meeting, New Orleans, LA, United States, Feb. 11-15, 2001
 (2001), 31-36. Editor(s): Hryn, John N. Publisher: Minerals, Metals &
 Materials Society, Warrendale, Pa.
 CODEN: 69BDNU
 DT Conference
 LA English
 CC 56-2 (Nonferrous Metals and Alloys)
 AB One of the major issues for users and potential users of **magnesium**
 is **magnesium melt** protection. This is particularly so
 with the use of sulfur hexafluoride coming under increased scrutiny, as it
 is now recognized as a very potent greenhouse gas. It has been
 demonstrated that sulfur hexafluoride prevents **molten**
magnesium from oxidation by adding fluorine into the unprotective
 magnesium oxide surface film. The hydrofluorocarbon gas
 1,1,1,2-tetrafluoroethane (HFC-134a) has been recently developed for
magnesium melt protection by the Cooperative Research
 Center for Cast Metals Manufacturing (CAST) in Australia. This gas has a
 global
 warming potential 18 times lower than sulfur hexafluoride and has been
 shown to be an effective replacement for SF6 in many applications.
 HFC-134a has also been shown to have some unique properties resulting in

more possible applications for **magnesium melt** protection than sulfur hexafluoride. Recent experience with HFC-134a in laboratory and plant trials is discussed as well as the global implications of replacing sulfur hexafluoride with hydrofluorocarbons.

ST **magnesium alloy melting** hydrofluorocarbon protection

IT Casting of metals
(hydrofluorocarbons as replacement for sulfur hexafluoride for protection of **molten magnesium** alloys)

IT Cast alloys
RL: PEP (Physical, engineering or chemical process); PRP (Properties); PROC (Process)
(magnesium alloys; hydrofluorocarbons as replacement for sulfur hexafluoride for protection of **molten magnesium** alloys)

IT 7439-95-4P, Magnesium, preparation
RL: IMF (Industrial manufacture); PREP (Preparation)
(hydrofluorocarbons as replacement for sulfur hexafluoride for protection of **molten magnesium** alloys)

IT **811-97-2**, 1,1,1,2-Tetrafluoroethane
RL: NUU (Other use, unclassified); PEP (Physical, engineering or chemical process); PROC (Process); USES (Uses)
(hydrofluorocarbons as replacement for sulfur hexafluoride for protection of **molten magnesium** alloys)

IT 135860-09-2, Az91d
RL: PEP (Physical, engineering or chemical process); PRP (Properties); PROC (Process)
(hydrofluorocarbons as replacement for sulfur hexafluoride for protection of **molten magnesium** alloys)

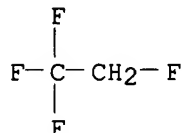
RE.CNT 20 THERE ARE 20 CITED REFERENCES AVAILABLE FOR THIS RECORD

RE

- (1) Anon; Private communications with Solvay und Fluor GmbH
- (2) Bartos, S; Magnesium Technology 2000 2000, P83
- (3) Cashion, S; PhD thesis, The University of Queensland 1998
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- (5) Couling, S; 36th Annual World Conference on Magnesium 1979
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- (7) Erickson, S; Foundry Management & Technology 1998, V126(6), P38
- (8) Erickson, S; International Magnesium Association Technical Committee Report
- (9) Fruehling, J; PhD thesis, University of Michigan 1970
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IT **811-97-2**, 1,1,1,2-Tetrafluoroethane
RL: NUU (Other use, unclassified); PEP (Physical, engineering or chemical process); PROC (Process); USES (Uses)
(hydrofluorocarbons as replacement for sulfur hexafluoride for

protection of molten magnesium alloys)
 RN 811-97-2 HCAPLUS
 CN Ethane, 1,1,1,2-tetrafluoro- (6CI, 7CI, 8CI, 9CI) (CA INDEX NAME)



L54 ANSWER 5 OF 9 METADEX COPYRIGHT 2004 CSA on STN
 AN 2001(12):51-2066 METADEX
 TI **Hydrofluorocarbons** as a replacement for SF6.
 AU Ricketts, N.J. (CSIRO); Cashion, S.P. (University of Queensland)
 SO Magnesium Industry (Sept. 2001) 2, (5), 26-31, Graphs
 ISSN: 1590-959X
 DT Journal
 CY Italy
 LA English
 AB One of the major issues for users and potential users of **magnesium** is **magnesium melt** protection. This is particularly so with the use of sulphur hexafluoride coming under increased scrutiny, as it is now recognized as a very potent greenhouse gas. It has been demonstrated that sulphur hexafluoride prevents **molten magnesium** from oxidation by adding fluorine into the unprotective magnesium oxide surface film. The **hydrofluorocarbon** gas 1,1,1,2-tetrafluoroethane (HFC-134a) has been recently developed for **magnesium melt** protection by the Cooperative Research Centre for Cast Metals Manufacturing (CAST) in Australia. This gas has a global warming potential 18 times lower than sulphur hexafluoride and has been shown to be an effective replacement for SF6 in many applications. HFC-134a has also been shown to have some unique properties resulting in more possible applications for **magnesium melt** protection than sulphur hexafluoride.
 CC 51 Foundry
 CT Journal Article; **Magnesium: Melting**; Oxidation resistance: Processing effects; Combustion: Processing effects; Fluorides: Materials selection; Halocarbons; Sulfur compounds: Materials substitution
 ET F*S; SF6; S cp; cp; F cp

L54 ANSWER 6 OF 9 HCAPLUS COPYRIGHT 2004 ACS on STN
 AN 2000:772535 HCAPLUS
 DN 133:312399
 ED Entered STN: 03 Nov 2000
 TI Cover gas mixtures for protection of reactive **Mg** or **Mg**
 -alloy **melt**
 IN Ricketts, Nigel Jeffrie; Frost, Malcolm Timothy; Cashion, Simon Paul;
 Korn, Craig John; Baker, Phillip Wilmott
 PA Cast Centre Pty Ltd, Australia
 SO PCT Int. Appl., 20 pp.
 CODEN: PIXXD2
 DT Patent
 LA English
 IC ICM B22D021-04
 ICS B22D021-02; A62D001-02; C22B026-22; C22B004-02; C22C001-02;
 C22C023-00

applicants

CC 56-2 (Nonferrous Metals and Alloys)
 Section cross-reference(s): 50

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 2000064614	A1	20001102	WO 2000-AU393	20000428
	W: AE, AG, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, CA, CH, CN, CR, CU, CZ, DE, DK, DM, DZ, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, TZ, UA, UG, US, UZ, VN, YU, ZA, ZW, AM, AZ, BY, KG, KZ, MD, RU, TJ, TM				
	RW: GH, GM, KE, LS, MW, SD, SL, SZ, TZ, UG, ZW, AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE, BF, BJ, CF, CG, CI, CM, GA, GN, GW, ML, MR, NE, SN, TD, TG				
	BR 2000010137	A	20020122	BR 2000-10137	20000428
	TR 200103096	T2	20020422	TR 2001-200103096	20000428
	EP 1204499	A1	20020515	EP 2000-920274	20000428
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO, MK, CY, AL				
	JP 2002541999	T2	20021210	JP 2000-613596	20000428
	AU 766844	B2	20031023	AU 2000-40930	20000428
	NZ 515084	A	20031031	NZ 2000-515084	20000428
	TW 500805	B	20020901	TW 2000-89108130	20000504
	NO 2001005264	A	20011221	NO 2001-5264	20011026
	ZA 2001008862	A	20021028	ZA 2001-8862	20011026
	BG 106138	A	20020628	BG 2001-106138	20011123
PRAI	AU 1999-15	A	19990428		
	WO 2000-AU393	W	20000428		

CLASS

PATENT NO.	CLASS	PATENT FAMILY CLASSIFICATION CODES
WO 2000064614	ICM	B22D021-04
	ICS	B22D021-02; A62D001-02; C22B026-22; C22B004-02; C22C001-02; C22C023-00

AB The cover-gas blend for protection of the **Mg** or **Mg**-alloy **molten** bath (especially in casting) contains: (a) hydrofluorocarbon type inhibiting gas, typically as difluoromethane or heptafluoropropane at <1% by volume; and (b) suitable carrier gas, especially dry

air, CO₂, Ar, and/or N₂. The component gases are selected to have no O₃-depletion potential, as well as a low Global-Warming potential. The cover-gas blends are suitable for preventing oxidation of **Mg** alloy **melt** for ingot casting, as well as for extinguishing of accidental fires on the molten bath. The protective mixture suitable for preventing surface oxidation on the cast ingots of AZ91 Mg-Al alloy contains 0.04% by volume of 1,1,1,2-tetrafluoroethane in dry air, vs. surface dross and discoloration on the ingots solidified under the similar amount of SF₆ in dry air.

ST **magnesium** alloy **melt** casting cover gas protection; hydrofluorocarbon cover gas magnesium fire prevention

IT Hydrocarbons, uses

RL: TEM (Technical or engineered material use); USES (Uses) (fluoro, gas mixts. with; cover gas mixts. with hydrofluorocarbons for protection of reactive **Mg-melt** bath)

IT Safety

(foundry, for Mg alloy casting; cover gas mixts. with hydrofluorocarbons for protection of reactive **Mg-alloy melt**)

IT Casting of metals
(of Mg alloys; cover gas mixts. with hydrofluorocarbons for protection of reactive **Mg-alloy melt**)

IT Fire
(prevention, on **Mg-alloy melt**; cover gas mixts. with hydrofluorocarbons for protection of reactive **Mg-alloy melt**)

IT 12634-54-7, AZ91
RL: PEP (Physical, engineering or chemical process); PROC (Process) (casting of, surface protection in; cover gas mixts. with hydrofluorocarbons for protection of reactive **Mg-melt bath**)

IT 124-38-9, Carbon dioxide, uses 7440-37-1, Argon, uses 7727-37-9, Nitrogen, uses
RL: MOA (Modifier or additive use); USES (Uses) (gas mixts. with; cover gas mixts. with hydrofluorocarbons for protection of reactive **Mg-melt bath**)

IT 75-10-5, Difluoromethane 75-37-6 354-33-6, Pentafluoroethane 811-97-2, 1,1,1,2-Tetrafluoroethane 33660-75-2, Heptafluoropropane 138495-42-8, HFC-43-10 163702-05-4, HFE-7200 219484-64-7, HFE-7100
RL: TEM (Technical or engineered material use); USES (Uses) (gas mixts. with; cover gas mixts. with hydrofluorocarbons for protection of reactive **Mg-melt bath**)

IT 7439-95-4, Magnesium, uses
RL: TEM (Technical or engineered material use); USES (Uses) (**molten**; cover gas mixts. with hydrofluorocarbons for protection of reactive **Mg-melt bath**)

RE.CNT 6 THERE ARE 6 CITED REFERENCES AVAILABLE FOR THIS RECORD

RE

- (1) Dougherty; US 5115868 A 1992 HCAPLUS
- (2) Great Lakes Chemical Corporation; WO 9102564 A 1991
- (3) Li; US 5855647 A 1999 HCAPLUS
- (4) Minnesota Mining And Manufacturing Company; WO 9622129 A 1996 HCAPLUS
- (5) Radzilowski; US 4214899 A 1980 HCAPLUS
- (6) Tokai Rika Denki Kk; JP 08143985 1996 HCAPLUS

IT 75-10-5, Difluoromethane 75-37-6 354-33-6, Pentafluoroethane 811-97-2, 1,1,1,2-Tetrafluoroethane 33660-75-2, Heptafluoropropane 138495-42-8, HFC-43-10 163702-05-4, HFE-7200 219484-64-7, HFE-7100
RL: TEM (Technical or engineered material use); USES (Uses) (gas mixts. with; cover gas mixts. with hydrofluorocarbons for protection of reactive **Mg-melt bath**)

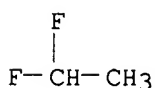
RN 75-10-5 HCAPLUS

CN Methane, difluoro- (8CI, 9CI) (CA INDEX NAME)

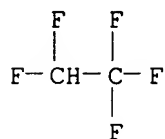
F-CH₂-F

RN 75-37-6 HCAPLUS

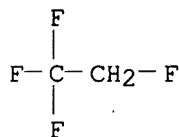
CN Ethane, 1,1-difluoro- (6CI, 8CI, 9CI) (CA INDEX NAME)



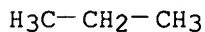
RN 354-33-6 HCAPLUS
 CN Ethane, pentafluoro- (6CI, 7CI, 8CI, 9CI) (CA INDEX NAME)



RN 811-97-2 HCAPLUS
 CN Ethane, 1,1,1,2-tetrafluoro- (6CI, 7CI, 8CI, 9CI) (CA INDEX NAME)

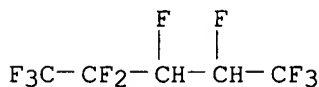


RN 33660-75-2 HCAPLUS
 CN Propane, heptafluoro- (6CI, 7CI, 8CI, 9CI) (CA INDEX NAME)

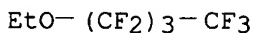


7 (D1- F)

RN 138495-42-8 HCAPLUS
 CN Pentane, 1,1,1,2,2,3,4,5,5,5-decafluoro- (9CI) (CA INDEX NAME)



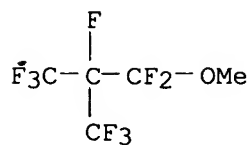
RN 163702-05-4 HCAPLUS
 CN Butane, 1-ethoxy-1,1,2,2,3,3,4,4,4-nonafluoro- (9CI) (CA INDEX NAME)



RN 219484-64-7 HCAPLUS
 CN Butane, 1,1,1,2,2,3,3,4,4-nonafluoro-4-methoxy-, mixt. with
 2-(difluoromethoxymethyl)-1,1,1,2,3,3,3-heptafluoropropane (9CI) (CA
 INDEX NAME)

CM 1

CRN 163702-08-7
 CMF C5 H3 F9 O



CM 2

CRN 163702-07-6
CMF C5 H3 F9 O

MeO-(CF₂)₃-CF₃

IT 7439-95-4, **Magnesium**, uses
RL: TEM (Technical or engineered material use); USES (Uses)
(molten; cover gas mixts. with hydrofluorocarbons for
protection of reactive **Mg-melt** bath)
RN 7439-95-4 HCAPLUS
CN Magnesium (8CI, 9CI) (CA INDEX NAME)

Mg

L54 ANSWER 7 OF 9 HCAPLUS COPYRIGHT 2004 ACS on STN
AN 1999:811058 HCAPLUS
DN 132:54870
ED Entered STN: 24 Dec 1999
TI Pharmaceutical formulations for aerosols with two or more active
substances
IN McNamara, Daniel P.; Destefano, George A.
PA Boehringer Ingelheim Pharmaceuticals, Inc., USA
SO PCT Int. Appl., 18 pp.
CODEN: PIXXD2
DT Patent
LA English
IC ICM A61K009-12
ICS A61K031-135; A61K031-46
CC 63-6 (Pharmaceuticals)
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 9965464	A1	19991223	WO 1999-US12785	19990608
	W: AE, AU, BG, BR, BY, CA, CN, CZ, EE, HR, HU, ID, IL, IN, JP, KR, KZ, LT, LV, MX, NO, NZ, PL, RO, RU, SG, SI, SK, TR, UA, UZ, VN, YU, ZA				
	RW: AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE				
	DE 19827178	A1	20000427	DE 1998-19827178	19980618
	DE 19842963	A1	20000323	DE 1998-19842963	19980919
	CA 2335065	AA	19991223	CA 1999-2335065	19990608
	AU 9945521	A1	20000105	AU 1999-45521	19990608
	AU 759222	B2	20030410		

BR 9911351	A	20010313	BR 1999-11351	19990608
EP 1087750	A1	20010404	EP 1999-928458	19990608
EP 1087750	B1	20031112		
R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO				
TR 200003721	T2	20010621	TR 2000-200003721	19990608
ZA 200007581	A	20020402	ZA 2000-7581	19990608
EE 200000759	A	20020415	EE 2000-759	19990608
NZ 509418	A	20030630	NZ 1999-509418	19990608
JP 2003522102	T2	20030722	JP 2000-554344	19990608
AT 253896	E	20031115	AT 1999-928458	19990608
RU 2218152	C2	20031210	RU 2001-101898	19990608
PT 1087750	T	20040227	PT 1999-928458	19990608
ES 2211108	T3	20040701	ES 1999-928458	19990608
TW 528606	B	20030421	TW 1999-88110010	19990615
BG 105033	A	20010928	BG 2000-105033	20001207
NO 2000006318	A	20010130	NO 2000-6318	20001212
HR 2000000867	A1	20011031	HR 2000-867	20001215
PRAI DE 1998-19827178	A	19980618		
DE 1998-19842963	A	19980919		
WO 1999-US12785	W	19990608		

CLASS

PATENT NO.	CLASS	PATENT FAMILY CLASSIFICATION CODES
WO 9965464	ICM	A61K009-12
	ICS	A61K031-135; A61K031-46
AB	The present invention relates to new pharmaceutical formulations for aerosols with at least two or more active substances for administration by inhalation or by nasal route. Specifically, the invention relates to pharmaceutical preps. for propellant-driven metered dose aerosols using a fluorohydrocarbon (HFC) as propellant, which contain a combination of active substance of two or more active substances, wherein at least one active substance is present in dissolved form together with at least one other active substance in the form of suspended particles. An aerosol composition contained ipratropium bromide 37 mg, salbutamol sulfate 210.5 mg, citric acid 4 mg, liquefied TG 134a 89.96 g, ethanol 10.03 g, and oleic acid 50 mg.	
ST	fluorohydrocarbon propellant aerosol ipratropium salbutamol	
IT	Fatty acids, biological studies	
	RL: THU (Therapeutic use); BIOL (Biological study); USES (Uses) (C5-20; aerosols containing two or more active substances and fluorohydrocarbon propellant)	
IT	Carbohydrates, biological studies	
	Glycerides, biological studies	
	Lecithins	
	RL: THU (Therapeutic use); BIOL (Biological study); USES (Uses) (aerosols containing two or more active substances and fluorohydrocarbon propellant)	
IT	Drug delivery systems	
	(aerosols; aerosols containing two or more active substances and fluorohydrocarbon propellant)	
IT	Quaternary ammonium compounds, biological studies	
	RL: THU (Therapeutic use); BIOL (Biological study); USES (Uses) (alkylbenzyltrimethyl, chlorides; aerosols containing two or more active substances and fluorohydrocarbon propellant)	
IT	Fatty acids, biological studies	
	RL: THU (Therapeutic use); BIOL (Biological study); USES (Uses) (esters, C5-20; aerosols containing two or more active substances and fluorohydrocarbon propellant)	

IT Alcohols, biological studies
 RL: THU (Therapeutic use); BIOL (Biological study); USES (Uses)
 (fatty, C5-20; aerosols containing two or more active substances and
 fluorohydrocarbon propellant)

IT Hydrocarbons, biological studies
 RL: THU (Therapeutic use); BIOL (Biological study); USES (Uses)
 (fluoro; aerosols containing two or more active substances and
 fluorohydrocarbon propellant)

IT Drug delivery systems
 (inhalants; aerosols containing two or more active substances and
 fluorohydrocarbon propellant)

IT Drug delivery systems
 (nasal sprays; aerosols containing two or more active substances and
 fluorohydrocarbon propellant)

IT 64-17-5, Ethanol, biological studies 112-80-1, Oleic acid, biological
 studies 431-89-0, TG 227 586-06-1, Orciprenaline 811-97-2,
 TG 134a 1338-43-8, Sorbitan monooleate 1944-12-3, Fenoterol
 hydrobromide 3385-03-3, Flunisolide 4419-39-0, Beclomethasone
 12441-09-7D, Sorbitan, esters 13392-18-2, Fenoterol 16110-51-3,
 Cromoglycic acid 18559-94-9, Salbutamol 22254-24-6, Ipratropium
 bromide 23031-25-6, Terbutalin 26266-58-0, Sorbitan trioleate
 29116-98-1, Sorbitan dioleate 30286-75-0, Oxitropium bromide
 51022-70-9, Albuterol sulfate 51333-22-3, Budesonide 54063-54-6,
 Reproterol 60205-81-4, Ipratropium 69049-73-6, Nedocromil
 89365-50-4, Salmeterol 90566-53-3, Fluticasone 148430-28-8
 RL: THU (Therapeutic use); BIOL (Biological study); USES (Uses)
 (aerosols containing two or more active substances and fluorohydrocarbon
 propellant)

IT 50-81-7, Ascorbic acid, biological studies 60-00-4, EDTA, biological
 studies 77-92-9, Citric acid, biological studies 7647-01-0,
 Hydrochloric acid, biological studies 7664-38-2, Phosphoric acid,
 biological studies 7664-93-9, Sulfuric acid, biological studies
 7697-37-2, Nitric acid, biological studies
 RL: THU (Therapeutic use); BIOL (Biological study); USES (Uses)
 (as stabilizer; aerosols containing two or more active substances and
 fluorohydrocarbon propellant)

RE.CNT 5 THERE ARE 5 CITED REFERENCES AVAILABLE FOR THIS RECORD

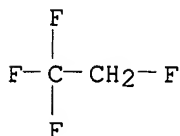
RE

(1) Bell, A; WO 9801147 A 1998 HCAPLUS
 (2) Henry, R; US 5589156 A 1996 HCAPLUS
 (3) Jager, P; WO 9413262 A 1994 HCAPLUS
 (4) McNamara, D; US 5603918 A 1997 HCAPLUS
 (5) Riker Laboratories Inc; EP 0499344 A 1992 HCAPLUS

IT 811-97-2, TG 134a
 RL: THU (Therapeutic use); BIOL (Biological study); USES (Uses)
 (aerosols containing two or more active substances and fluorohydrocarbon
 propellant)

RN 811-97-2 HCAPLUS

CN Ethane, 1,1,1,2-tetrafluoro- (6CI, 7CI, 8CI, 9CI) (CA INDEX NAME)



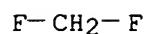
L54 ANSWER 8 OF 9 HCAPLUS COPYRIGHT 2004 ACS on STN
 AN 1995:422771 HCAPLUS
 DN 122:165333
 ED Entered STN: 18 Mar 1995
 TI Parting compounds containing fluorocarbons which contain no chlorine and may contain hydrogen
 IN Buchwald, Hans; Hellmann, Joachim; Raschkowski, Boleslaus
 PA Solvay Fluor und Derivate GmbH, Germany
 SO Ger. Offen., 6 pp.
 CODEN: GWXXBX
 DT Patent
 LA German
 IC ICM C10M111-02
 ICS B29C033-60; C03B040-02
 ICA C09K015-30; C09K015-22; C09K015-18; C09K003-30; B01F017-00; B22C001-14
 ICI C10M111-02, C10M105-52, C10M105-34, C10M105-02, C10M105-76, C10M105-24, C10M105-72, C10M107-04, C10M107-18, C10N040-36
 CC 51-11 (Fossil Fuels, Derivatives, and Related Products)
 Section cross-reference(s): 38, 57
 FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	DE 4321288	A1	19950105	DE 1993-4321288	19930626
	US 5476603	A	19951219	US 1994-262930	19940621
	FI 9403069	A	19941227	FI 1994-3069	19940623
	JP 07150188	A2	19950613	JP 1994-141905	19940623
PRAI	DE 1993-4321288	A	19930626		

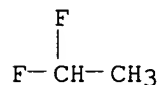
CLASS

	PATENT NO.	CLASS	PATENT FAMILY CLASSIFICATION CODES
	DE 4321288	ICM	C10M111-02
		ICS	B29C033-60; C03B040-02
		ICA	C09K015-30; C09K015-22; C09K015-18; C09K003-30; B01F017-00; B22C001-14
		ICI	C10M111-02, C10M105-52, C10M105-34, C10M105-02, C10M105-76, C10M105-24, C10M105-72, C10M107-04, C10M107-18, C10N040-36
US	5476603	ECLA	B29C033/60; B29C033/62; B29C033/64
AB	The parting compds. are pressure-liquefied and contain a fluorocarbon chosen from among R134a, R143a, R125, R32, R143, R134, R23, R152a, R14, and R116, a parting agent (e.g., wax, paraffins, silicones, metallic soaps, fats, oils, polymers, or inorg. powders), and optionally solvents (e.g., a gasoline fraction b.p. 100-140°).		
ST	parting agent fluorocarbon		
IT	Solvents (pressure-liquefied fluorocarbon-containing parting compns.)		
IT	Fats and Glyceridic oils Gasoline Oils Paraffin waxes and Hydrocarbon waxes, uses Polymers, uses Siloxanes and Silicones, uses Soaps Waxes and Waxy substances RL: TEM (Technical or engineered material use); USES (Uses) (pressure-liquefied fluorocarbon-containing parting compns.)		
IT	Hydrocarbons, uses RL: TEM (Technical or engineered material use); USES (Uses) (fluoro, pressure-liquefied fluorocarbon-containing parting compns.)		

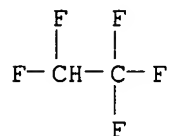
IT 75-10-5, Difluoromethane 75-37-6, R152a 75-46-7, R23
 75-73-0, R14 76-16-4, R116 354-33-6, R125 359-35-3, R134
 420-46-2, R143a 430-66-0, R143 557-04-0, **Magnesium** stearate
 811-97-2, R134a 7429-90-5D, Aluminum, soaps 7439-92-1D, Lead,
 soaps 7439-95-4D, **Magnesium**, soaps 7440-66-6D, Zinc,
 soaps 7440-70-2D, Calcium, soaps
 RL: TEM (Technical or engineered material use); USES (Uses)
 (pressure-liquefied fluorocarbon-containing parting compns.)
 IT 75-10-5, Difluoromethane 75-37-6, R152a 354-33-6
 , R125 811-97-2, R134a 7439-95-4D, **Magnesium**
 , soaps
 RL: TEM (Technical or engineered material use); USES (Uses)
 (pressure-liquefied fluorocarbon-containing parting compns.)
 RN 75-10-5 HCAPLUS
 CN Methane, difluoro- (8CI, 9CI) (CA INDEX NAME)



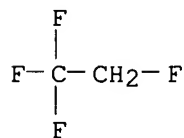
RN 75-37-6 HCAPLUS
 CN Ethane, 1,1-difluoro- (6CI, 8CI, 9CI) (CA INDEX NAME)



RN 354-33-6 HCAPLUS
 CN Ethane, pentafluoro- (6CI, 7CI, 8CI, 9CI) (CA INDEX NAME)



RN 811-97-2 HCAPLUS
 CN Ethane, 1,1,1,2-tetrafluoro- (6CI, 7CI, 8CI, 9CI) (CA INDEX NAME)



RN 7439-95-4 HCAPLUS
 CN Magnesium (8CI, 9CI) (CA INDEX NAME)

Mg

L54 ANSWER 9 OF 9 HCAPLUS COPYRIGHT 2004 ACS on STN

AN 1954:64093 HCAPLUS

DN 48:64093

OREF 48:11316h-i,11317a-g

ED Entered STN: 22 Apr 2001

TI Pyrolyses of the salts of the perfluoro carboxylic acids

AU La Zerte, J. D.; Hals, L. J.; Reid, T. S.; Smith, G. H.

CS Minnesota Mining & Manufg. Co., St. Paul

SO Journal of the American Chemical Society (1953), 75, 4525-8

CODEN: JACSAT; ISSN: 0002-7863

DT Journal

LA Unavailable

CC 10 (Organic Chemistry)

AB The thermal decomposition of a number of salts of the straight-chain perfluoro acids has been investigated. From the Na salts, terminally unsatd. perfluoroolefins were prepared in yields ranging from 65 to 100%. The reaction is represented by the equation $C_nF_{2n+1}CF_2CF_2CO_2Na \rightarrow C_nF_{2n+1}CF:CF_2 + CO_2 + NaF$. Salts of other metals of the groups I, II, and III of the periodic table gave varying yields of olefins. $C_3F_7CO_2Ag$ (I) and $C_7F_{15}CO_2Ag$ (II) decomposed to give C_6F_{14} and $C_{14}F_{30}$, resp. C_2F_4 was formed when a mixture of CF_3CO_2Na and NaOH was heated. A series of fluorocarbon hydrides, $C_nF_{2n+1}H$, was prepared by heating the salts of perfluoro acids in $(CH_2OH)_2$. The NH_4 , Li, Na, K, Ca, Sr, and Ba salts of the perfluoro acids were all prepared by neutralizing an aqueous solution of

the

acid with a solution of the hydroxide. $(C_3F_7CO_2)_2Mg$ and $(C_3F_7CO_2)_2Pb$ were obtained from aqueous $C_3F_7CO_2H$ (III) and the metal oxides at slightly above 25° ; both salts were hygroscopic; the vacuum-dried Pb salt was further dried by azeotropic distillation with CCl_4 . I and II were prepared by treating freshly prepared Ag_2O with the dilute aqueous acids. $(C_3F_7CO_2)_2Cu$ was obtained by passing dry air into a mixture of finely divided Cu powder and excess III at 120° . $(C_3F_7CO_2)_3Al$ was prepared by the method of Hood and Ihde (C.A. 44, 7228i) from $AlCl_3$ and excess III in the presence of $(C_3F_7CO_2)_2O$ (IV) at 100° . The purity of the salts had a great influence on the decomposition reaction. In the presence of an inorg. base, the pyrolysis of the salts gave products contaminated with fluorocarbon monohydrides; to avoid this, the pH of the salt solns. was adjusted to pH 5-7. H_2O vapors in the pyrolysis zone also led to the formation of H-containing compds. The pyrolyses were carried out, in general, in Pyrex flasks; the rate of the decomposition was controlled by varying the

temperature; the

resulting volatile products were passed through 2 scrubbers containing 15% KOH, dried over P_2O_5 , and condensed in a cold trap. The thermal stabilities of some salts of III were determined by heating small weighed samples 0.5 hr. at $20-5^\circ$ intervals until almost complete decomposition was obtained; the temperature at which 20% decomposition was obtained (given)

was for

the following salts: NH_4 185° , K 200° , Na 235° , Ba

275° , Sr 275° , Ag 295° ; and for $(CF_3)_2CFCO_2Na$

185° . The Na salts of higher straight-chain perfluoro acids

underwent 20% decomposition at $240-50^\circ$, and $C_4F_9CO_2K$ at $175-80^\circ$.

The Na and Ba salts of CF_3CO_2H gave CF_3COF and $(CF_3CO)_2O$; the same

products were obtained from the Li and Ca salts. The pyrolysis of

CF_3CO_2Na in the presence of solid NaOH proceeded at about 270°

exothermically to give C_2F_4 , along with some CF_3COF and CHF_3 ; the min.

yield of C_2F_4 was 32% in better than 98% purity; 1% by weight of Pr_3N was

always added to the C_2F_4 to prevent the explosive polymerization of the

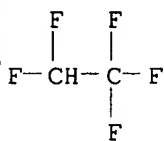
monomer. The following salts of III were pyrolyzed and the pyrolysis

products determined (the decomposition temperature, % yield $CF_3CF:CF_2$, and the

other

fluorinated products formed given): Li, 240-50°, 20, C₃F₇COF (V), IV, III; K, 215-35°, 98, -; Mg, 275-300°, <5, high-boiling liquid; Ca, 275-300°, <10, V, IV, III; Sr, 275-85°, 25, V, III; Ba, 265-75°, 78, -; Pb, 300-5°, <10, V, IV, some III; Cu, trace, V, unidentified product; Al, 250°, <5, V, III, C₂F₆; NH₄, 180-200°, 0, CF₃CF₂CF₂H; Ag, 300-20°, 45, C₆F₁₄. C₄F₉CO₂K (1681 g.) and 907 g. (CH₂OH)₂ heated 5 hrs. at 170-90° gave 1169 g. cold-trap condensate which on fractionation yielded 1017 g. (84%) CF₃(CF₂)₃H, b₇₄₀ 14°, λ_{maximum} 3015 cm.⁻¹ (C-H). Similarly were prepared from the Na salts of the appropriate perfluoro acids the following hydrides CF₃(CF₂)_nH (VI) (n, % yield, b.p./740 mm. given): 1, 98, -50°; 2, 97, -16°; 4, 80, 46°; 6, 60, 94°, n_{25D} 1.2690. C₅F₁₁CO₂Na (210 g.); prepared in 93% yield by neutralizing C₅F₁₁CO₂H with aqueous NaOH, pyrolyzed at about 250° yielded 141 g. (90%) C₃F₇CF:CF₂, b. 28-9.0°, n_{25D} 1.2571, λ_{maximum} 1795 cm.⁻¹ Similarly were prepared the following olefins from the appropriate Na salts (compound, % yield, b.p., and n_{15D} given): C₂F₄, 90, -74°, -; CF₃CF:CF₂, 97, -29°, -; C₂F₅CF:CF₂, 91, 1°, -; C₅F₁₁CF:CF₂, 86, 81, 1.2782; C₇F₁₅CF:CF₂, 65, 123°, 1.2868. The infrared absorption spectra of VI with n = 2, 3, 4, and 6 all showed C-H absorption in the range 2940-2990 cm.⁻¹.

- IT Spectra
(of hydrocarbons (fluorinated))
- IT Pyrolysis
(of perfluoro carboxylic acid salts)
- IT Salts
(of perfluoro carboxylic acids, pyrolysis of)
- IT Octanoic acid, pentadecafluoro-, silver salt
Propionic acid, 2,3,3,3-tetrafluoro-2-(trifluoromethyl)-, sodium salt
(pyrolysis of)
- IT 76-05-1, Acetic acid, trifluoro-
(and derivs.)
- IT 375-22-4, Butyric acid, heptafluoro-
(and derivs., pyrolysis of)
- IT 116-14-3, Ethylene, tetrafluoro-
(formation of, from CF₃Co₂Na and NaOH)
- IT 75-46-7, Fluoroform
(formation of, from Na trifluoroacetate)
- IT 76-16-4, Ethane, hexafluoro-
(formation of, in (C₃F₇CO₂)₃Al pyrolysis)
- IT 116-15-4, Propene, hexafluoro-
(formation of, in pyrolysis of heptafluorobutyric acid salts)
- IT 307-62-0, Tetradecane, triacontafluoro- 354-33-6, Ethane, pentafluoro- 354-34-7, Acetyl fluoride, trifluoro- 355-42-0, Hexane, tetradecafluoro- 355-63-5, 1-Heptene, tetradecafluoro- 357-26-6, 1-Butene, octafluoro- 375-17-7, Butane, 1,1,1,2,2,3,3,4,4-nonafluoro- 375-61-1, Pentane, 1,1,1,2,2,3,3,4,4,5,5-undecafluoro- 375-83-7, Heptane, 1,1,1,2,2,3,3,4,4,5,5,6,6,7,7-pentadecafluoro- 376-22-7, 1-Nonene, octadecafluoro- 376-87-4, 1-Pentene, decafluoro- 2252-84-8, Propane, 1,1,1,2,2,3,3-heptafluoro-
(preparation of)
- IT 336-23-2, Valeric acid, nonafluoro-, potassium salt 2923-26-4, Hexanoic acid, undecafluoro-, sodium salt
(pyrolysis of)
- IT 354-33-6, Ethane, pentafluoro-
(preparation of)
- RN 354-33-6 HCAPLUS
- CN Ethane, pentafluoro- (6CI, 7CI, 8CI, 9CI) (CA INDEX NAME)



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